

### Nombrar los complejos IV: (Solución a formular complejos III)

- 1  $[\text{Cu}(\text{NH}_3)_4]^{+2}$
- 2  $[\text{CrCl}_2(\text{H}_2\text{O})_4]^{+1}$
- 3  $[\text{Ag}(\text{NH}_3)_2]^{+1}$
- 4  $[\text{Fe}(\text{CN})_6]^{-4}$
- 5  $[\text{Al}(\text{H}_2\text{O})_6]_2(\text{SO}_4)_3$
- 6  $[\text{Cu}(\text{CN})_2]^{-1}$
- 7  $[\text{Hg}(\text{SO}_3)_2]^{-2}$
- 8  $[\text{OsBr}(\text{NH}_3)_3]$
- 9  $\text{Cs}_2[\text{TaF}_7]$
- 10  $[\text{CoCl}_3(\text{NH}_3)_3]$
- 11  $[\text{Co}(\text{NH}_3)_6]\text{Cl}_3$
- 12  $[\text{Co}(\text{NH}_3)_4]\text{SO}_4$
- 13  $\text{NH}_4[\text{Au}(\text{CN})_2]$
- 14  $[\text{Cu}(\text{NH}_3)_4]\text{Cl}_2$
- 15  $[\text{Cu}(\text{NH}_3)_4]\text{Cl}$
- 16  $[\text{Cr}(\text{H}_2\text{O})_6]\text{Cl}_3$
- 17  $\text{Li}_4[\text{Co}(\text{CN})_4]$
- 18  $\text{K}[\text{PtCl}_5\text{NH}_3]$
- 19  $[\text{PtCl}_2(\text{NH}_3)_4]\text{Cl}_2$
- 20  $\text{K}_2[\text{Pt}(\text{NO}_2)_4]$
- 21  $\text{Rb}_2[\text{Zn}(\text{OH})_4]$
- 22  $\text{K}_4[\text{Fe}(\text{CN})_6]$
- 23  $(\text{NH}_4)_2[\text{PtCl}_6]$
- 24  $[\text{Co}(\text{H}_2\text{O})(\text{NH}_3)_5]_2(\text{SO}_4)_3$
- 25  $[\text{CoCl}_3(\text{NH}_3)_3]$
- 26  $[\text{Cr}(\text{NO}_2)_3(\text{NH}_3)_3]$
- 27  $[\text{CuCl}_2(\text{NH}_3)_2]$
- 28  $[\text{Ni}(\text{CO})_4]$
- 29  $[\text{CoCl}_2(\text{NH}_3)_4]$
- 30  $[\text{FeCl}_3(\text{NH}_3)_2(\text{H}_2\text{O})]$
- 31  $[\text{Cr}(\text{NH}_3)_4(\text{CN})_2]^{+1}$
- 32  $\text{K}_4[\text{Fe}(\text{CN})_6]$
- 33  $[\text{Cr}(\text{NH}_3)_5\text{Cl}]\text{Cl}_2$
- 34  $[\text{Cr}(\text{H}_2\text{O})_2\text{F}_2(\text{O}_2)_2]^{-3}$
- 35  $\text{Na}[\text{Co}(\text{CN})_3(\text{CO})_2(\text{NO})]$
- 36  $[\text{CrCl}_3(\text{NH}_3)_3]$
- 37  $[\text{FeBr}_2(\text{CO})_4]$
- 38  $[\text{Ag}(\text{NH}_3)_2]^{+1}$

II Nombrar los complejos:

- 1  $[\text{Cd}(\text{NH}_3)_4]^{+2}$
- 2  $[\text{CrCl}_2(\text{NH}_3)_4]\text{Cl}$
- 3  $[\text{Co}(\text{H}_2\text{O})_6]\text{Cl}_2$
- 4  $[\text{CrCl}(\text{NH}_3)_5]\text{Cl}_2$
- 5  $\text{K}_3[\text{Fe}(\text{CN})_6]$
- 6  $\text{K}_4[\text{Fe}(\text{CN})_6]$
- 7  $[\text{Cr}(\text{CrO}_4)(\text{H}_2\text{O})_5]\text{NO}_3$
- 8  $[\text{CoH}(\text{CO})_4]$
- 9  $[\text{Fe}(\text{H}_2\text{O})_6]^{+3}$
- 10  $[\text{Ag}(\text{NH}_3)_2]^{+1}$
- 11  $\text{Li}_2[\text{Pt}(\text{NO}_2)_4]$
- 12  $[\text{Co}(\text{SO}_4)(\text{H}_2\text{O})_4]_2\text{SO}_4$
- 13  $[\text{FeF}_6]^{-3}$
- 14  $[\text{Ir}(\text{AsO}_4)(\text{H}_2\text{O})_5]$
- 15  $[\text{AlH}_4]^{-1}$
- 16  $[\text{Fe}(\text{H}_2\text{O})_6]^{+3}$
- 17  $\text{Li}[\text{AlH}_4]$
- 18  $[\text{Cu}(\text{NH}_3)_4]_3(\text{PO}_4)_2$
- 19  $\text{K}[\text{Au}(\text{OH})_4]$
- 20  $[\text{Co}(\text{NH}_3)_6]^{+3}$
- 21  $[\text{Co}(\text{NH}_3)_6][\text{Cr}(\text{CO}_3)_3]$
- 22  $[\text{CrBr}(\text{NH}_3)_5]^{+2}$
- 23  $[\text{CrBr}(\text{NH}_3)_5]\text{Cl}_2$
- 24  $[\text{Cu}(\text{NH}_3)_4]^{+2}$
- 25  $[\text{PtF}_4(\text{NH}_3)_2]$
- 26  $[\text{CrCl}_2(\text{H}_2\text{O})_4]\text{NO}_2$
- 27  $[\text{Zn}(\text{H}_2\text{O})_6]\text{SO}_4$
- 28  $\text{H}_2[\text{PbCl}_6]$
- 29  $[\text{Cd}(\text{NH}_3)_4](\text{NO}_3)_2$
- 30  $\text{K}_2[\text{Pt}(\text{CN})_4]$
- 31  $\text{Mg}_2[\text{Fe}(\text{CN})_6]$
- 32  $\text{Ni}[\text{Fe}(\text{CN})_4]$
- 33  $\text{Ca}[\text{Hg}(\text{SO}_4)_2]$
- 34  $\text{Ca}[\text{Ni}(\text{OH})_4(\text{H}_2\text{O})_2]$
- 35  $[\text{Co}_2(\text{CO})_8]$
- 36  $[\text{Co}(\text{H}_2\text{O})_2(\text{NH}_3)_4]\text{Cl}_3$
- 37  $[\text{Fe}(\text{CO}_3)_4\text{Cl}_2]^{-7}$
- 38  $[\text{Cr}(\text{H}_2\text{O})_2(\text{NO}_2)_3]$
- 39  $[\text{Cu}(\text{NH}_3)_2(\text{OH})_2]$

### III Nombrar los complejos:

- 1  $[\text{Cu}(\text{NH}_3)_4] \text{SO}_4$
- 2  $\text{K}_3 [\text{Fe}(\text{CN})_6]$
- 3  $[\text{PtCl}_6]^{-2}$
- 4  $[\text{Ag}(\text{NH}_3)_2]^{-1}$
- 5  $[\text{CrCl}(\text{NH}_3)_5]^{+2}$
- 6  $[\text{Co}(\text{NH}_3)_6] \text{Cl}_3$
- 7  $[\text{Cd}(\text{NH}_3)_4] (\text{NO}_3)_2$
- 8  $[\text{CuCl}_4]^{-2}$
- 9  $[\text{HgBr}_2\text{O}]^{-2}$
- 10  $[\text{Cu}(\text{CN})_2]^{-1}$
- 11  $[\text{Cr}(\text{OH})(\text{H}_2\text{O})_5]^{+2}$
- 12  $[\text{OsBr}(\text{NH}_3)_5] \text{Br}_2$
- 13  $\text{K}_2 [\text{Ru}(\text{NO}_2)_4 \text{OH NO}]$
- 14  $\text{K}_2 [\text{RuCl}_5 \text{NO}]$
- 15  $\text{H} [\text{AuCl}_4]$
- 16  $\text{Cs} [\text{Au}(\text{OH})_4]$
- 17  $[\text{CoOH}(\text{NH}_3)_5] \text{SO}_3$
- 18  $[\text{FeCl}_2(\text{NH}_3)_4] \text{Cl}$
- 19  $[\text{Cu}(\text{CN})_4]^{-2}$
- 20  $[\text{Cd}(\text{NH}_3)_4]^{+2}$
- 21  $[\text{OsBr}_2(\text{NH}_3)_4]^{+1}$
- 22  $\text{Rb} [\text{Cu}(\text{CN})_2]$
- 23  $[\text{CrOH}(\text{H}_2\text{O})_5] (\text{NO}_3)_2$
- 24  $[\text{Cu}(\text{OH})_2(\text{H}_2\text{O})_4]^{-1}$
- 25  $[\text{NiBr}_2(\text{H}_2\text{O})_4]^{+1}$
- 26  $\text{Cr} [\text{PtCl}_6]$
- 27  $\text{Na}_2 [\text{PtBr}_4 \text{Cl}_2]$
- 28  $[\text{Zn}(\text{H}_2\text{O})_6]^{+2}$
- 29  $[\text{Cd}(\text{NH}_3)_4]^{-2}$
- 30  $\text{Cu}_3 [\text{Co}(\text{NO}_2)_6]_2$
- 31  $\text{Al} [\text{CrF}_4 \text{O}]$
- 32  $[\text{Sn}(\text{OH})_4]^{-2}$
- 33  $\text{Na} [\text{AlH}_4]$
- 34  $[\text{Fe}(\text{NH}_3)_4 \text{Cl}_2]^{+1}$
- 35  $[\text{Co}(\text{NH}_3)_6] \text{F}_3$
- 36  $\text{Al}[\text{CoCl}_6]$
- 37  $\text{K}_2[\text{PtCl}_4]$
- 38  $[\text{PtCl}_3 \text{NO}_2 (\text{NH}_3)_2]$
- 39  $\text{Al}[\text{Co}(\text{NO}_2)_6]$