

## Nombrar los complejos -II- : (Solución a formular complejos I )

- 1  $[\text{Pt Cl}_2(\text{NH}_3)_2]$
- 2  $[\text{Ni}(\text{NH}_3)_6](\text{NO}_3)_2$
- 3  $[\text{Fe}(\text{H}_2\text{O})_2(\text{NH}_3)_4]^{+2}$
- 4  $[\text{Ir}(\text{AsO}_4)(\text{H}_2\text{O})_5]$
- 5  $[\text{Al}(\text{OH})_2(\text{H}_2\text{O})_4]_2 \text{SO}_4$
- 6  $[\text{Pt Cl Br}(\text{CH}_3 - \text{NH}_2)_4]^{+2}$
- 7  $\text{Cs}_2 [\text{Ta F}_7]$
- 8  $\text{NH}_4[\text{AlH}_4]$
- 9  $[\text{Cr O}_2 (\text{O}_2) (\text{CN})_2 \text{NH}_3]^{-5}$
- 10  $[\text{Pt Cl}_3 \text{NO}_2]^{-2}$
- 11  $[\text{Fe}(\text{CO})_5]$
- 12  $\text{Rb}_2[\text{Zn}(\text{OH})_4]$
- 13  $[\text{Cr F Cl O}_2 (\text{H}_2\text{O})_2]^{-3}$
- 14  $[\text{Rh}(\text{NO}_3)(\text{NH}_3)_5] \text{Br}_2$
- 15  $[\text{Pd Br Cl F}]$
- 16  $\text{Sn}_3[\text{Nb F}_6 \text{O}]_2$
- 17  $[\text{Co}(\text{NO}_3)_6]^{-3}$
- 18  $\text{K}_3[\text{Fe}(\text{CN})_5(\text{NO})]$
- 19  $\text{Ag}[\text{Pt Br Cl}(\text{NO}_2)(\text{NH}_3)]$
- 20  $\text{Ca}_2[\text{Mo Cl}_8]$
- 21  $[\text{Cr}(\text{CrO}_4)(\text{H}_2\text{O})_4]_2 \text{Cr}_2\text{O}_7$
- 22  $[\text{Zn}(\text{OH})_3(\text{H}_2\text{O})]^{-1}$
- 23  $[\text{Cr}(\text{NH}_3)_6][\text{Co}(\text{CO}_3)_3]$
- 24  $[\text{Mo}(\text{SCN})(\text{NH}_3)_5] \text{Cl}_2$
- 25  $[\text{Ru}(\text{NH}_3)_5(\text{N}_2)] \text{Br}_2$
- 26  $\text{Cd}[\text{Hg Cl}_2 (\text{CN})_2]$
- 27  $\text{Na} [\text{Co}(\text{CO})_4]$
- 28  $\text{K}_4[\text{Fe}(\text{CN})_6]$
- 29  $[\text{Cr}(\text{OH})(\text{NH}_3)_2(\text{H}_2\text{O})_3] \text{NO}_3$
- 30  $[\text{Co}(\text{NH}_3)_6][\text{Cr CO}_3]_3$
- 31  $[\text{Cr}(\text{CO})_6]$
- 32  $[\text{Cu}(\text{NH}_3)_4] (\text{NO}_3^-)_2$
- 33  $\text{Al}_2 [\text{Pt}(\text{CN})_4]_3$
- 34  $\text{K}_2[\text{Co}(\text{NO}_2)_5(\text{NH}_3)]$