

I Nombrar los complejos IV: (Solución a formular complejos III)

- 1 $[\text{Cu}(\text{NH}_3)_4]^{+2}$
- 2 $[\text{CrCl}_2(\text{H}_2\text{O})_4]^{+1}$
- 3 $[\text{Ag}(\text{NH}_3)_2]^{+1}$
- 4 $[\text{Fe}(\text{CN})_6]^{-4}$
- 5 $[\text{Al}(\text{H}_2\text{O})_6]_2(\text{SO}_4)_3$
- 6 $[\text{Cu}(\text{CN})_2]^{-1}$
- 7 $[\text{Hg}(\text{SO}_3)_2]^{-2}$
- 8 $[\text{OsBr}(\text{NH}_3)_3]$
- 9 $\text{Cs}_2[\text{TaF}_7]$
- 10 $[\text{CoCl}_3(\text{NH}_3)_3]$
- 11 $[\text{Co}(\text{NH}_3)_6]\text{Cl}_3$
- 12 $[\text{Co}(\text{NH}_3)_4]\text{SO}_4$
- 13 $\text{NH}_4[\text{Au}(\text{CN})_2]$
- 14 $[\text{Cu}(\text{NH}_3)_4]\text{Cl}_2$
- 15 $[\text{Cu}(\text{NH}_3)_4]\text{Cl}$
- 16 $[\text{Cr}(\text{H}_2\text{O})_6]\text{Cl}_3$
- 17 $\text{Li}_4[\text{Co}(\text{CN})_4]$
- 18 $\text{K}[\text{PtCl}_5\text{NH}_3]$
- 19 $[\text{PtCl}_2(\text{NH}_3)_4]\text{Cl}_2$
- 20 $\text{K}_2[\text{Pt}(\text{NO}_2)_4]$
- 21 $\text{Rb}_2[\text{Zn}(\text{OH})_4]$
- 22 $\text{K}_4[\text{Fe}(\text{CN})_6]$
- 23 $(\text{NH}_4)_2[\text{PtCl}_6]$
- 24 $[\text{Co}(\text{H}_2\text{O})(\text{NH}_3)_5]_2(\text{SO}_4)_3$
- 25 $[\text{CoCl}_3(\text{NH}_3)_3]$
- 26 $[\text{Cr}(\text{NO}_2)_3(\text{NH}_3)_3]$
- 27 $[\text{CuCl}_2(\text{NH}_3)_2]$
- 28 $[\text{Ni}(\text{CO})_4]$
- 29 $[\text{CoCl}_2(\text{NH}_3)_4]$
- 30 $[\text{FeCl}_3(\text{NH}_3)_2(\text{H}_2\text{O})]$
- 31 $[\text{Cr}(\text{NH}_3)_4(\text{CN})_2]^{+1}$
- 32 $\text{K}_4[\text{Fe}(\text{CN})_6]$
- 33 $[\text{Cr}(\text{NH}_3)_5\text{Cl}]\text{Cl}_2$
- 34 $[\text{Cr}(\text{H}_2\text{O})_2\text{F}_2(\text{O}_2)_2]^{-3}$
- 35 $\text{Na}[\text{Co}(\text{CN})_3(\text{CO})_2(\text{NO})]$
- 36 $[\text{CrCl}_3(\text{NH}_3)_3]$
- 37 $[\text{FeBr}_2(\text{CO})_4]$
- 38 $[\text{Ag}(\text{NH}_3)_2]^{+1}$

II Nombrar los complejos:

- 1 $[\text{Cd}(\text{NH}_3)_4]^{+2}$
- 2 $[\text{CrCl}_2(\text{NH}_3)_4]\text{Cl}$
- 3 $[\text{Co}(\text{H}_2\text{O})_6]\text{Cl}_2$
- 4 $[\text{CrCl}(\text{NH}_3)_5]\text{Cl}_2$
- 5 $\text{K}_3[\text{Fe}(\text{CN})_6]$
- 6 $\text{K}_4[\text{Fe}(\text{CN})_6]$
- 7 $[\text{Cr}(\text{CrO}_4)(\text{H}_2\text{O})_5]\text{NO}_3$
- 8 $[\text{CoH}(\text{CO})_4]$
- 9 $[\text{Fe}(\text{H}_2\text{O})_6]^{+3}$
- 10 $[\text{Ag}(\text{NH}_3)_2]^{+1}$
- 11 $\text{Li}_2[\text{Pt}(\text{NO}_2)_4]$
- 12 $[\text{Co}(\text{SO}_4)(\text{H}_2\text{O})_4]_2\text{SO}_4$
- 13 $[\text{FeF}_6]^{-3}$
- 14 $[\text{Ir}(\text{AsO}_4)(\text{H}_2\text{O})_5]$
- 15 $[\text{AlH}_4]^{-1}$
- 16 $[\text{Fe}(\text{H}_2\text{O})_6]^{+3}$
- 17 $\text{Li}[\text{AlH}_4]$
- 18 $[\text{Cu}(\text{NH}_3)_4]_3(\text{PO}_4)_2$
- 19 $\text{K}[\text{Au}(\text{OH})_4]$
- 20 $[\text{Co}(\text{NH}_3)_6]^{+3}$
- 21 $[\text{Co}(\text{NH}_3)_6][\text{Cr}(\text{CO}_3)_3]$
- 22 $[\text{CrBr}(\text{NH}_3)_5]^{+2}$
- 23 $[\text{CrBr}(\text{NH}_3)_5]\text{Cl}_2$
- 24 $[\text{Cu}(\text{NH}_3)_4]^{+2}$
- 25 $[\text{PtF}_4(\text{NH}_3)_2]$
- 26 $[\text{CrCl}_2(\text{H}_2\text{O})_4]\text{NO}_2$
- 27 $[\text{Zn}(\text{H}_2\text{O})_6]\text{SO}_4$
- 28 $\text{H}_2[\text{PbCl}_6]$
- 29 $[\text{Cd}(\text{NH}_3)_4](\text{NO}_3)_2$
- 30 $\text{K}_2[\text{Pt}(\text{CN})_4]$
- 31 $\text{Mg}_2[\text{Fe}(\text{CN})_6]$
- 32 $\text{Ni}[\text{Fe}(\text{CN})_4]$
- 33 $\text{Ca}[\text{Hg}(\text{SO}_4)_2]$
- 34 $\text{Ca}[\text{Ni}(\text{OH})_4(\text{H}_2\text{O})_2]$
- 35 $[\text{Co}_2(\text{CO})_8]$
- 36 $[\text{Co}(\text{H}_2\text{O})_2(\text{NH}_3)_4]\text{Cl}_3$
- 37 $[\text{Fe}(\text{CO}_3)_4\text{Cl}_2]^{-7}$
- 38 $[\text{Cr}(\text{H}_2\text{O})_2(\text{NO}_2)_3]$
- 39 $[\text{Cu}(\text{NH}_3)_2(\text{OH})_2]$

III Nombrar los complejos:

- 1 $[\text{Cu}(\text{NH}_3)_4] \text{SO}_4$
- 2 $\text{K}_3 [\text{Fe}(\text{CN})_6]$
- 3 $[\text{Pt} \text{Cl}_6]^{-2}$
- 4 $[\text{Ag}(\text{NH}_3)_2]^{-1}$
- 5 $[\text{Cr} \text{Cl}(\text{NH}_3)_5]^{+2}$
- 6 $[\text{Co}(\text{NH}_3)_6] \text{Cl}_3$
- 7 $[\text{Cd}(\text{NH}_3)_4] (\text{NO}_3)_2$
- 8 $[\text{Cu} \text{Cl}_4]^{-2}$
- 9 $[\text{Hg} \text{Br}_2 \text{O}]^{-2}$
- 10 $[\text{Cu}(\text{CN})_2]^{-1}$
- 11 $[\text{Cr}(\text{OH})(\text{H}_2\text{O})_5]^{+2}$
- 12 $[\text{Os} \text{Br}(\text{NH}_3)_5] \text{Br}_2$
- 13 $\text{K}_2 [\text{Ru}(\text{NO}_2)_4 \text{OH} \text{NO}]$
- 14 $\text{K}_2 [\text{Ru} \text{Cl}_5 \text{NO}]$
- 15 $\text{H} [\text{Au} \text{Cl}_4]$
- 16 $\text{Cs} [\text{Au}(\text{OH})_4]$
- 17 $[\text{Co} \text{OH}(\text{NH}_3)_5] \text{SO}_3$
- 18 $[\text{Fe} \text{Cl}_2(\text{NH}_3)_4] \text{Cl}$
- 19 $[\text{Cu}(\text{CN})_4]^{-2}$
- 20 $[\text{Cd}(\text{NH}_3)_4]^{+2}$
- 21 $[\text{Os} \text{Br}_2(\text{NH}_3)_4]^{+1}$
- 22 $\text{Rb} [\text{Cu}(\text{CN})_2]$
- 23 $[\text{Cr} \text{OH}(\text{H}_2\text{O})_5] (\text{NO}_3)_2$
- 24 $[\text{Cu}(\text{OH})_2(\text{H}_2\text{O})_4]^{-1}$
- 25 $[\text{Ni} \text{Br}_2(\text{H}_2\text{O})_4]^{+1}$
- 26 $\text{Cr} [\text{Pt} \text{Cl}_6]$
- 27 $\text{Na}_2 [\text{Pt} \text{Br}_4 \text{Cl}_2]$
- 28 $[\text{Zn}(\text{H}_2\text{O})_6]^{+2}$
- 29 $[\text{Cd}(\text{NH}_3)_4]^{-2}$
- 30 $\text{Cu}_3 [\text{Co}(\text{NO}_2)_6]_2$
- 31 $\text{Al} [\text{Cr} \text{F}_4 \text{O}]$
- 32 $[\text{Sn}(\text{OH})_4]^{-2}$
- 33 $\text{Na} [\text{Al} \text{H}_4]$
- 34 $[\text{Fe}(\text{NH}_3)_4 \text{Cl}_2]^{+1}$
- 35 $[\text{Co}(\text{NH}_3)_6] \text{F}_3$
- 36 $\text{Al}[\text{Co} \text{Cl}_6]$
- 37 $\text{K}_2[\text{Pt} \text{Cl}_4]$
- 38 $[\text{Pt} \text{Cl}_3 \text{NO}_2(\text{NH}_3)_2]$
- 39 $\text{Al}[\text{Co}(\text{NO}_2)_6]$